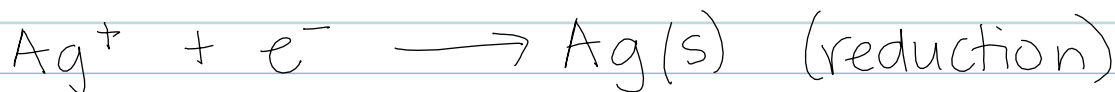
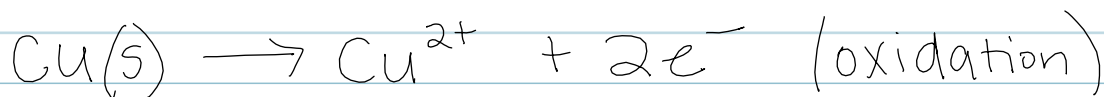
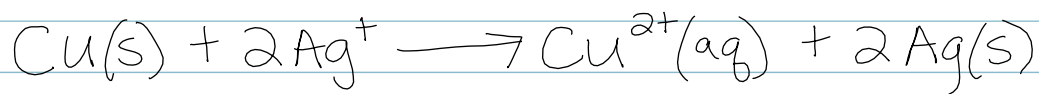


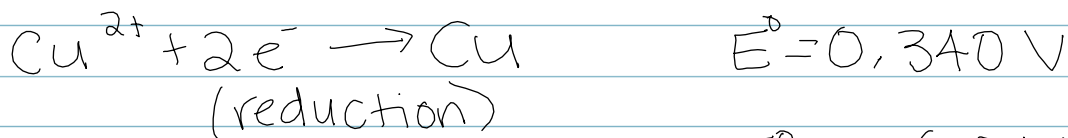
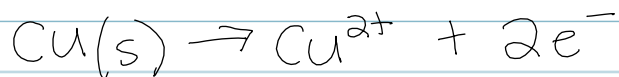
CH 1010 Cell Voltage / Cell Potential



Standard Reduction Potentials at 25°C

<u>Half RXN</u>	<u>E° (V)</u>
$\text{Ag}^+ + e^- \longrightarrow \text{Ag}(s)$	0.800
$\text{Cu}^{2+} + 2e^- \longrightarrow \text{Cu}(s)$	0.340

$$E^\circ = 0.800\text{V}$$



$$E^\circ = -0.340\text{V}$$

$$E^\circ_{\text{cell}} = E^\circ_{\text{cathode}} - E^\circ_{\text{anode}}$$

(Reduction) (oxidation)

$$E^\circ_{\text{cell}} = 0.800\text{V} - (-0.340\text{V})$$
$$= 1.140\text{V}$$